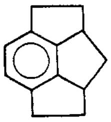
On page 3326, Table III, the data for H-8β should read as follows and footnote *b* should be introduced at the bottom of the table.

Proton	No. of examples	Av chem shift, δ	Coupling constant, J (cps)	Exceptions
H-8β	16	2.9 ± 0.15	8β,8α (12 ± 2)^b 8β,7β (8 ± 0.5)^b 8β,7α (4.3 ± 0.7)^b	IX (2.67)

[b] Known only for the compounds in Table II.

Crowded Benzenes. VI. The Strain Energy in *o*-Di-*t*-Butylbenzenes [*J. Am. Chem. Soc.*, **89**, 5389 (1967)]. By EDWARD M. ARNETT, JOSEPH C. SANDA, J. M. BOLLINGER, and M. BARBER. Department of Chemistry, University of Pittsburgh, and Mellon Institute, Pittsburgh, Pennsylvania 15213.

On page 5399, structure V should be

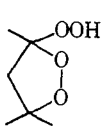


Deuterium Migration during the Acid-Catalyzed Dehydration of 6-Deuterio-5,6-dihydroxy-3-chloro-1,3-cyclohexadiene, a Nonenzymatic Model for the NIH Shift [*J. Am. Chem. Soc.*, **89**, 5488 (1967)]. By DONALD M. JERINA, JOHN W. DALY, and BERNHARD WITKOP. National Institutes of Health, Bethesda, Maryland.

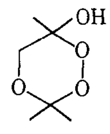
Structures XII and XVI on the flow sheet on page 5488 should be interchanged so that loss of a proton (deuteron) leads from XI to XII and from XV to XVI.

Ozonolysis. Evidence for Carbonyl Oxide Tautomerization and for 1,3-Dipolar Addition to Olefins [*J. Am. Chem. Soc.*, **89**, 5726 (1967)]. By PAUL R. STORY and JOHN R. BURGESS. Department of Chemistry, The University of Georgia, Athens, Georgia 30601.

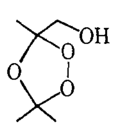
The structure assigned to compound **6** has been found to be incorrect. Although the evidence for carbonyl oxide tautomerization stands, the evidence for carbonyl oxide addition to an olefin is not valid. The structure **6** was assigned on the basis of several pieces of information, all but two of which were equally consistent with **6a** and **6b**, structures which could result from the reaction of carbonyl oxide with hydroxy acetone.



6



6a



6b

The structure **6** was chosen (and **6a** and **6b** eliminated) principally on the basis of hydride reduction which yielded only isopropyl alcohol. We now find, after more careful searching, that 1,3-propanediol is also present. Furthermore, Professor R. Criegee has pointed out to us that **6** has been synthesized by Rieche [A. Rieche, E. Schmitz, and E. Gundemann, *Chem. Ber.*, **93**, 2443 (1960)]. Our compound is not identical with Rieche's. Although we cannot decide unequivocally between **6a** and **6b** for the structure in question, the nmr evidence favors **6a**.

σ- and π-Bonding Effects in the Coordination of the Cyano Group [*J. Am. Chem. Soc.*, **89**, 6139 (1967)]. By KEITH F. PURCELL. Department of Chemistry, Kansas State University, Manhattan, Kansas 66502.

In Table II, the experimental value of f_z for N^{14} in CH_3NC should be <0.05 and not 0.5.

On page 6141, column 2, the entry for "Cl, Exptl" should read 0.72 (not 0.92).

Conformations of Cyclic Peptides. The Folding of Cyclic Dipeptides Containing an Aromatic Side Chain [*J. Am. Chem. Soc.*, **89**, 6193 (1967)]. By KENNETH D. KOPPLE and DAVID H. MARR. Department of Chemistry, Illinois Institute of Technology, Chicago, Illinois 60616.

On page 6195, all of the entropy change values given in Table II are too small by a factor of *R*. The average entropy change referred to in the discussion is thus 6–8 cal/mole deg.

Book Reviews

Mechanistic Organic Photochemistry. By DOUGLAS C. NECKERS, Department of Chemistry, Hope College, Holland, Mich. Reinhold Publishing Corp., 430 Park Ave., New York, N. Y. 1967. xii + 320 pp. 16 × 23.5 cm. $14.75.

The rapidly proliferating field of photochemical research has reached the point where there are four new textbooks, two new review series, and a new photochemical journal. Nearly all of these have appeared in the last four years, an observation which reflects the current popularity of photochemistry. This particular monograph, according to the author's preface, has been designed to fit the needs of undergraduate seniors and beginning graduate students who seek an introduction to the topic of organic photochemistry. In this respect, the author has done an admirable job.

The tone of the book is essentially descriptive and is quite comprehensive in its treatment of various classes of organic photoreactions. The first four chapters of the book provide background information and an introduction to photochemical reactions. After a brief introductory chapter (7 pp) dealing with the relationship of photochemistry to spectroscopy and kinetics, the next three chapters (47 pp) are devoted to a discussion of absorption processes, electronic transitions, the Franck–Condon principle, photophysical phenomena, excited-state multiplicity, and energy transfer. Although these three sections could have been much expanded, the present treatment provides both adequate orientation and entrée to the literature.

The longest part of the book, some 240 pages, is concerned with essentially all of the important aspects of photochemical reaction